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AMOUNT OF SUBSTANCE

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AMOUNT OF SUBSTANCE - Chemsheets

When an ionic substance dissolves in water, the positive and negative ions separate and become hydrated (they interact with water molecules rather than each other).

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Chemistry, you **MUST** be able to write a

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formula without a second thought. It is the single most essential skill for an A level chemist.

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AS level Chemistry - Amount of substance | Teaching Resources

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substance) ANS bdkw39.pdf - AMOUNT
OF SUBSTANCE TASK 1 Writing formulas
of ionic compounds 1 7 13 AgBr
Al(NO₃)₃ Ga(OH)₃ 2

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More Exam Questions on 1.2 Amount of
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Exercise 2 - solutions 1.2 Exercise 3 -
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empirical and molecular formulae 1.2
Exercise 5 - ionic equations 1.2 Exercise
6 - more complex calculations

1.2 Amount Of Substance - A-Level Chemistry

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40 then penalise this mark but allow
consequential M2 (0.0185) 0.018 with no
M r shown = 0

Amount of Substance Answers

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The amount of substance is sometimes referred to as the chemical amount. The mole (symbol: mol) is a unit of amount of substance in the International System of Units, defined (since 2019) by fixing the Avogadro constant at the given value. Historically, the mole was defined as the amount of substance in 12 grams of the carbon-12 isotope.

Amount of substance - Wikipedia

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- The amount of energy needed to make 1 g of a substance 1 °C (1 K) hotter is

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called the specific heat capacity (measured in $\text{J g}^{-1} \text{K}^{-1}$). • The following equation is then used to find the amount of heat energy give out (or absorbed). $q = m c \Delta\Delta\Delta\Delta T$ $q =$ heat energy given out (J) $m =$ mass of substance heated (g)

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Chemsheets AS 1037 Equation (write oxidation states) Redox reaction (/)
Dispropor-tionation (/ Species

REDOX REACTIONS OR NOT?

The molar mass of an element is found on the periodic table, and it is the element's atomic weight in grams/mole (g/mol). If the mass of a substance is known, the number of moles in the substance can be calculated. Converting the mass, in grams, of a substance to moles requires a conversion factor of (one mole of substance/molar mass of substance).

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